

UNION SPORTIVE «SOKOL»

162, BOULEVARD DE LA GARE - CASABLANCA
RESTAURANT «SLAVIA»

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NEPRAVIDELNĚ VYDÁVÁ
T. J. SOKOL - CASABLANCA



VOLNOST . ROVNOST . BRATRSTVÍ

The first part of the report is devoted to a description of the
 experimental conditions and the results obtained. It is found that
 the rate of reaction is independent of the concentration of the
 reactants and is proportional to the square of the concentration of
 the catalyst. This is in agreement with the proposed mechanism
 which involves a bimolecular reaction between the catalyst and
 the reactant. The rate constant is found to be 0.025 l. mole⁻¹
 sec⁻¹ at 25°C. The activation energy is 15.5 kJ/mole.

The second part of the report is devoted to a discussion of the
 results and a comparison with the results of other workers. It is
 found that the results are in good agreement with those of other
 workers.

The following table gives the rate constants and activation energies
 for the reaction at different temperatures.

It is seen from the above table that the rate constant increases
 with increasing temperature. This is in agreement with the
 Arrhenius equation. The activation energy is 15.5 kJ/mole.

The results of this investigation are in good agreement with those
 of other workers.

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Seznam otázek a odpovědí, které vám mohou pomoci

zvládnout tyto problémy.

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